

Name _____

Honors Chemistry
Naming compounds

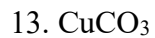
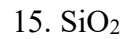
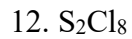
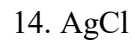
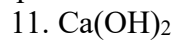
Determine the formula for the following compounds.

1. Cadmium (III)Nitrate
2. Calcium Carbonate
3. Manganese (V) Chlorate
4. Ammonium Sulfide.
5. Dinitrogen Monoxide.
6. Phosphorus trifluoride.

Complete the following conversions. You must first write the formula correctly!!

7. How many grams are in 3.5 mol of potassium arsenate?
8. How many moles are in 34 g of silver (I) carbonate?
9. How many grams are in 2.3 mol of chromium (II) phosphate?
10. How many moles are in 25 g of nitrogen triiodide?

Name the following the compounds. First determine whether they are ionic or covalent. If a compound has a metal or polyatomic ion it will form an ionic bond. Covalent compounds need prefixes. Ionic compounds do not.



Name the following **Acids**.

