Balance the following redox reactions

The following are in an acidic solution
$$1. \ Fe^{2+} \ + \ MnO_4^- \ \to \ Fe^{3+} \ + \ Mn^{2+}$$

$$2. \quad Sn^{2+} \quad + \qquad IO_3^{-} \qquad \qquad \rightarrow \qquad Sn^{4+} \quad + \qquad I^{-}$$

3.
$$S^{2-}$$
 + $NO_3^ \rightarrow$ S + NO

The following are in a basic solution 4. Zn + HgO \rightarrow

$$ZnO_2^{2-}$$
 + Hg

$$5. \quad Mn^{2+} \quad + \qquad \quad H_2O_2 \qquad \qquad \rightarrow \qquad \quad MnO_2 \qquad \qquad + \qquad \quad H_2O$$

6.
$$CrO_2 + ClO^- \rightarrow CrO_4^{2-} + Cl^-$$

$$\rightarrow$$
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