

Name _____

Review WS

1. What is the pH, pOH and $[H_3O^+]$ for a solution with a $[OH^-]$ of $2.34 \times 10^{-4} M$?

pH =

pOH =

$[H_3O^+]$ =

2. What is the pH, $[OH^-]$ and $[H_3O^+]$ for a solution with a pOH of 5.24?

pH =

$[OH^-]$ =

$[H_3O^+]$ =

3. What is the pH, pOH and $[OH^-]$ for a solution with a $[H_3O^+]$ of 3.8×10^{-9} ?

pH =

pOH =

$[OH^-]$ =

4. What is the pOH, $[OH^-]$ and $[H_3O^+]$ for a solution with a pH of 4.27?

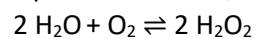
pOH =

$[OH^-]$ =

$[H_3O^+]$ =

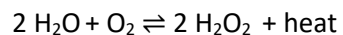
Write the stress shift and final for the following situations

5. For the equilibrium below, what will happen if more oxygen is added?

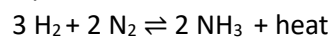


Write the stress shift and final for the following situations

6. For the equilibrium below, what will happen if the system is heated?



7. For the equilibrium below, what will happen if the amount of NH_3 is decreased?



Write the equilibrium constant, K, expression for the following equilibria

