

Ions with Lewis Dot structures

First things first!

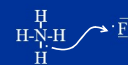
- Determine whether the bond will be ionic or covalent!
- A nonmetal and a metal is **ionic** if they are both nonmetals it is **covalent**, hydrogen is not a metal.
- To draw an ionic bond remove the electron(s) from the cation and add them to the anion. Show the charges on each ion.



For Polyatomic Ions...

- The polyatomic ion is a **covalently bonded** structure with an **ionic bond** to another substance.

- NH_4F

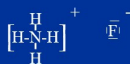


That goes to Fluoride
It has an extra electron

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Polyatomic Ions we need to know

Ammonium

What happens

What it looks like when done



Hypochlorite

What happens

What it looks like when done



Hydroxide

What happens

What it looks like when done



Nitrite

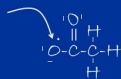
What happens

What it looks like when done

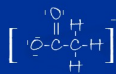


Acetate

What happens



What it looks like when done



Draw the Lewis Dot Structures for

- CCl_4
- MgF_2
- NH_4ClO
- $\text{Ca}(\text{NO}_2)_2$
- NaCH_3COO
- $\text{Al}(\text{OH})_3$