

Examples

- Fluorine is added to 2 propene
- Ethanol is burned in oxygen
- Chlorine is added to propane
- Ethanoic acid is reacted with 1-butanol

Ch 7 Predicting molecular equations

- Decomposition Reactions pg 42
- **CHECK CHARGES OF IONS!!!!!!**
- Carbonates decompose into oxides and CO_2
- Chlorates decompose into chlorides and O_2
- Ammonium and a base makes ammonia and water
- Hydrogen peroxide decomposes into water and oxygen.
- A binary compound may break down to produce two elements.

Ch 7 Predicting molecular equations

- Synthesis Reactions pg 42
- Metals and nonmetals combine to form salts.
CHECK CHARGES OF IONS!!!!!!
- Metal Oxides and water form bases, bases have OH attached
- Nonmetal oxides and water form acids
- Metal oxides and nonmetal oxides form salts (you will make an oxyanion- just move one oxygen over)

Three that appear on the AP test often

- **Hydrogen peroxide decomposes into water and oxygen**
- **Ammonium hydroxide decomposes into ammonia and water**
- **Carbonic acid decomposes into carbon dioxide and water**

Examples

- Magnesium burns in oxygen
- Calcium reacts with chlorine gas
- Magnesium oxide reacts with water
- Carbon dioxide is bubbled through water
- Lithium oxide is added to carbon dioxide

Examples

- Titanium (IV) chlorate decomposes
- Copper (I) carbonate is heated
- Magnesium chloride is electrolyzed
- Carbonic acid is heated
- Hydrogen peroxide decomposes